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Single Replacement Reaction
 A single replacement reaction occurs when one element replaces another element in a compound.

$A + BC \rightarrow AC + B$

Examples:
 $K + NaCl \rightarrow KCl + Na$
 $Cl_2 + 2NaBr \rightarrow 2NaCl + Br_2$

Chemistry Objective 6.02
 Worksheet – Balancing Equations

Identifying and Balancing Chemical Equations

Directions: Identify each of the equations below as synthesis, decomposition, single replacement or double replacement.

1. $HgO \rightarrow Hg + O_2$	
2. $NaCl + AgNO_3 \rightarrow NaNO_3 + AgCl$	
3. $Mg + HCl \rightarrow MgCl_2 + H_2$	
4. $Zn + H_2SO_4 \rightarrow ZnSO_4 + H_2$	
5. $NaOH + HCl \rightarrow NaCl + H_2O$	
6. $Al_2(SO_4)_3 + Ca(OH)_2 \rightarrow Al(OH)_3 + CaSO_4$	
7. $H_2 + O_2 \rightarrow H_2O$	
8. $Cl_2 + NaBr \rightarrow NaCl + Br_2$	
9. $Zn + CuSO_4 \rightarrow ZnSO_4 + Cu$	
10. $KClO_3 \rightarrow KCl + O_2$	
11. $H_2O + Fe \rightarrow Fe_2O_3 + H_2$	
12. $Ca(OH)_2 + HNO_3 \rightarrow Ca(NO_3)_2 + H_2O$	
13. $Na_2O + CO_2 \rightarrow Na_2CO_3$	
14. $H_2 + N_2 \rightarrow NH_3$	

Name: _____ Date: _____ Period: 1 2 3 4 5 6 7

Types of Chemical Reaction Worksheet

Balance the reactions 1 to 6 and indicate which type of chemical reaction (synthesis, decomposition, single-displacement, double-displacement or combustion) is being represented:

- $2_ NaBr + 1_ Ca(OH)_2 \rightarrow 1_ CaBr_2 + 2_ NaOH$ Reaction Type : **Double replacement**
- $2_ NH_3 + 1_ H_2SO_4 \rightarrow 1_ (NH_4)_2SO_4$ Reaction Type : **Synthesis**
- $4 C_2H_6O + 27 O_2 \rightarrow 20 CO_2 + 18 H_2O$ Reaction Type : **Combustion**
- $3_ Pb + 2_ H_3PO_4 \rightarrow 3_ H_2 + 1_ Pb_2(PO_4)_3$ Reaction Type : **single replacement**
- $1_ Li_3N + 3_ NH_4NO_3 \rightarrow 3_ LiNO_3 + 1_ (NH_4)_3N$ Reaction Type : **Double replacement**
- $3_ HBr + 1_ Al(OH)_3 \rightarrow 3_ H_2O + 1_ AlBr_3$ Reaction Type : **double replacement**
- $1_ Na_3PO_4 + 3_ KOH \rightarrow 3_ NaOH + 1_ K_3PO_4$ Reaction Type **double replacement**
- $1_ MgCl_2 + 1_ Li_2CO_3 \rightarrow 1_ MgCO_3 + 2_ LiCl$ Reaction Type **double replacement**
- $1_ C_6H_{16} + 12_ O_2 \rightarrow 8CO_2 + 8_ H_2O$ Reaction Type **combustion**

Indicate which type of chemical reaction (synthesis, decomposition, single-displacement, double-displacement or combustion) is being represented in 7 to 20.

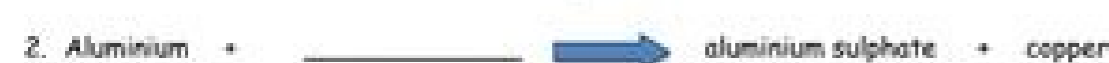
- $Pb + FeSO_4 \rightarrow PbSO_4 + Fe$ Reaction Type **single replacement**
- $CaCO_3 \rightarrow CaO + CO_2$ Reaction Type **decomposition**
- $P_4 + 3 O_2 \rightarrow 2 P_2O_3$ Reaction Type **synthesis**
- $2 RbNO_3 + BeF_2 \rightarrow Be(NO_3)_2 + 2 RbF$ Reaction Type **double replacement**
- $2 AgNO_3 + Cu \rightarrow Cu(NO_3)_2 + 2 Ag$ Reaction Type **single replacement**
- $C_3H_6O + 4 O_2 \rightarrow 3 CO_2 + 3 H_2O$ Reaction Type **combustion**
- $2 C_3H_5 + Fe \rightarrow Fe(C_3H_5)_2$ Reaction Type **synthesis**
- $SeCl_6 + O_2 \rightarrow SeO_2 + 3Cl_2$ Reaction Type **single replacement**
- $2 MgI_2 + Mn(SO_3)_2 \rightarrow 2 MgSO_3 + MnI_4$ Reaction Type **double replacement**
- $O_3 \rightarrow O + O_2$ Reaction Type **decomposition**
- $2 NO_2 \rightarrow 2 O_2 + N_2$ Reaction Type **decomposition**

Displacement reaction Word Equations

Decide whether or not the following reactions are possible (use your reactivity series to help you)



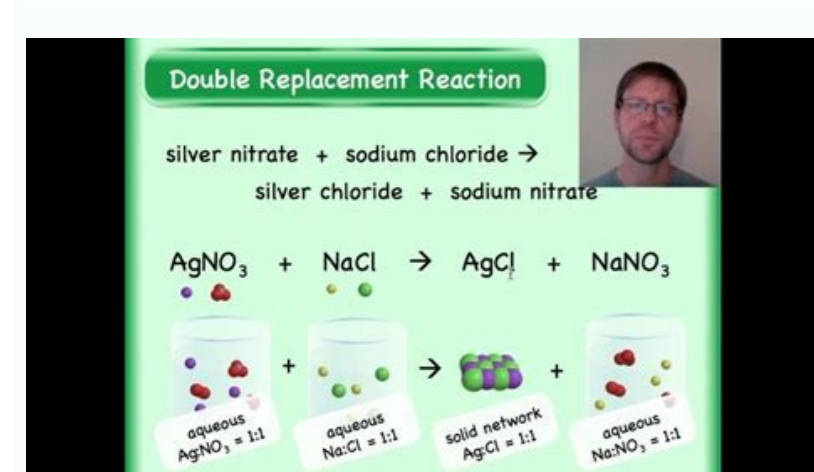
Complete the following word equations:



Extension questions:

Write your own word equations for the following reactions:

1. Zinc and copper sulphate
2. Magnesium and silver nitrate
3. Potassium and zinc nitrate
4. Gold and copper sulphate
5. Explain why aluminium and potassium nitrate will not react when mixed together.



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